

# Structural and Optical Characteristics of Cu Doped SnO<sub>2</sub> Composite Nanostructures

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## ABSTRACT

Copper doped tin oxide nanostructures were prepared by sol-gel method. Before that tin oxide and copper nanoparticles were synthesized using sol-gel and their structural and Optical properties were studied. The XRD analysis reveals that the CuO- SnO<sub>2</sub> nanoparticles. The crystallite sizes were also calculated. It has been proved from optical absorption measurements by FTIR that the nanometric size influences the energy band gap.

**Keywords: Cu, SnO<sub>2</sub>, Sol-gel, Optical, Structural**

## I. INTRODUCTION

Metal oxide nanoparticles have wide applications, primarily in the technology field, including their use as a semiconductor, electroluminescent or thermoelectric material, but they are also used in biomedical applications as drug delivery systems for treatment and diagnosis and in environmental decontamination applications. The classical methods for obtaining metal oxide nanoparticles are based on chemical and physical techniques that employ hazardous and expensive chemicals with high energy input and a negative effect on the environment. The production of metal oxide nanoparticles via biogenic synthesis has received increasing attention recently because it is a novel process for the development of engineered materials. Biogenic SnO<sub>2</sub> nanoparticles demonstrated antibacterial activity against *E. coli* and antioxidant properties, as assayed by scavenging the free radical of 2,2-diphenyl-1-picrylhydrazyl hydrate. These particles demonstrate biomedical applications. Cu concentration on the SnO<sub>2</sub> effects the sensibility which is not investigated as nano sensor in the literature.

## ► EXPERIMENTAL

### **Tin oxide, copper and copper doped tin oxides powders were prepared by sol-gel method:**

Before going to study about copper doped tin oxide nanostructures, copper and tin oxide nanoparticles are prepared and their structures are studied. The synthesis of Copper nanoparticles by mixing 4 gm of CuSO<sub>4</sub> to 0.9 gm of polyvinyl pyrrolidone after ½ hr stirring and keeping at 50°C for 30 min. Then 2.5-3 ml hydrazine hydrate was added dropwise. Colour changes observed from blue to copper. After that kept for stirring for 2 hrs. The synthesis of SnO<sub>2</sub> nanoparticles is started with one gm of stannous chloride dehydrate SnCl<sub>2</sub> .2H<sub>2</sub>O in a mixture of 11 ml deionized water and 5 ml carbanol in a beaker and stirred for one hour. Later 1 ml of HCl is added dropwise. After 1 hour, solution is fixed at PH=11 by dropwise to the above solution. This solution is

stirred for 2-3 hrs. at 80°C . The Cu doped SnO<sub>2</sub> nanopowder is achieved by adding dropwise solution of CuSO<sub>4</sub> while stirring in the above solution. All the prepared nanoparticles are annealed at 600°C.

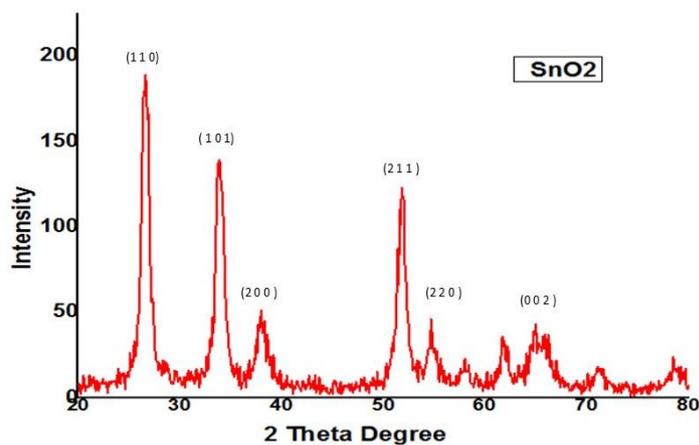
## II. RESULTS AND DISCUSSIONS

### 2.1 Characterization and computational details

The structural characterization was performed using X-ray diffractometer (type Bruker Advance-D8 with CuK $\alpha$  1.54 Å) .XRD revealed that powder formed is tin oxide which is an n-type semiconductor with a crystal structure of tetragonal lattice and a wide band gap of 3.6eV. This plays prominent role in surface acoustic wave gas sensors.

### Figures and Tables

Fig. XRD of SnO<sub>2</sub>



The miller indices (1 1 0), (1 0 1), (2 0 0), (2 1 1), (2 2 0) and (0 0 2) for pure SnO<sub>2</sub> XRD confirmed that the powder formed is FCC nano copper.

Fig: XRD of Copper nanopowder

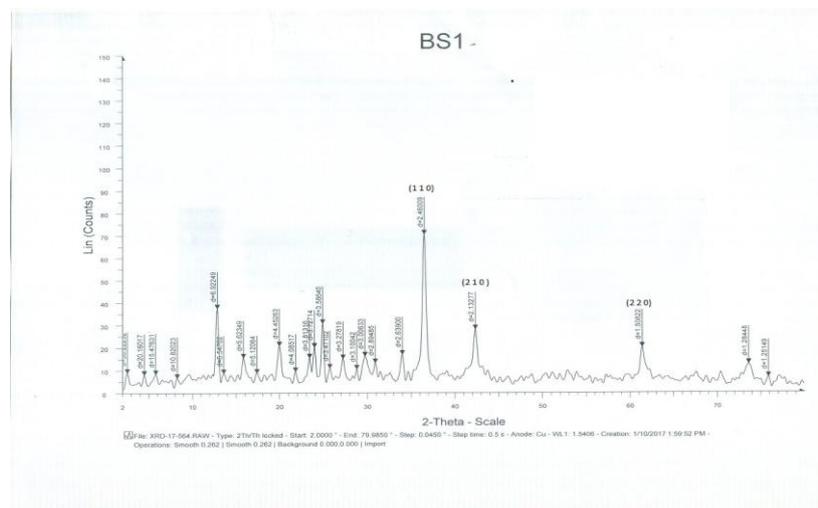
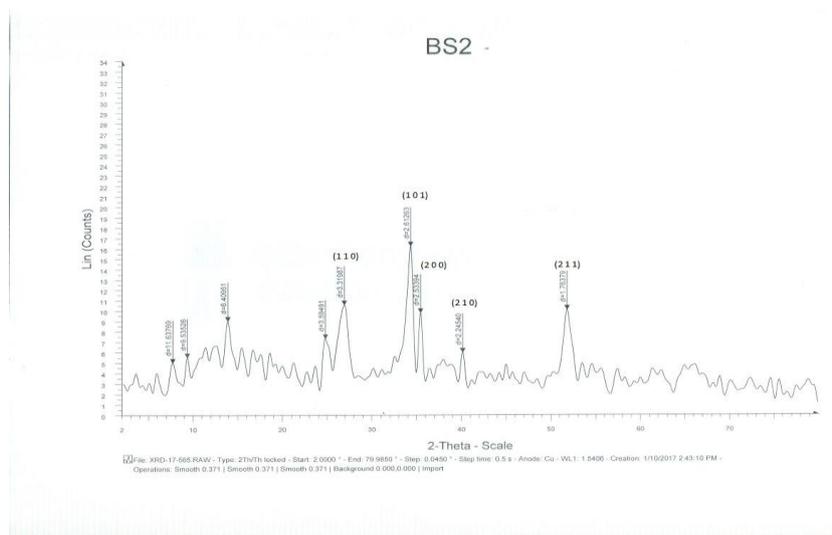


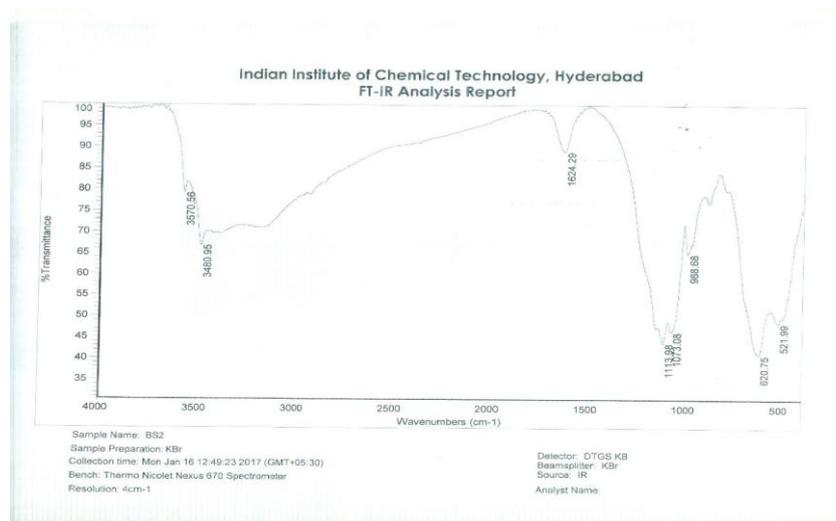
Fig: XRD of Cu doped SnO<sub>2</sub>



The XRD of Cu doped SnO<sub>2</sub> sample are shown above. Using the interplanar distance values, their corresponding miller indices have been identified and located in fig. It is clear that the Cu<sup>2+</sup> ions are acting as dopants in the SnO<sub>2</sub> structure. For tin oxide, suitable crystal face orientation is (1 1 0). Tin forms an interstitial bond with oxygen and exists either as SnO or SnO<sub>2</sub>, which has valency +2 or +4 respectively. During the process, the Cu substituted on Sn<sup>4+</sup> sites act as dopants and releasing excess electrons into the conduction band. Thus the band gap decreases with the addition of Cu. The mean crystallite size was calculated by Debye-Scherrer's formula  $D = K\lambda / \beta \cos\theta$  where K is a grain shaping factor 0.9,  $\lambda$  is the wavelength of the incident Cu-K $\alpha$  beam,  $\theta$  is a Bragg's angle and  $\beta$  is the full width at half maximum.

Table below gives the values of indices, d and crystallite sizes respectively.

Sample	(h k l)	d-spacing A <sup>o</sup>	crystallite size A <sup>o</sup>
Copper	(210)	2.13	25.32
SnO <sub>2</sub>	(1 1 0)	3.319	10.57
Cu doped SnO <sub>2</sub>	(1 1 0)	3.59	3.98



**Fig.FTIR of Cu doped SnO<sub>2</sub>**

From the above spectra it is clear that the variations indicate that Copper has incorporated in SnO<sub>2</sub> host lattice. The absorption peak at 521.99/cm and 3670/cm arises from the bending variation of Sn-O-Sn. At 3480.95/cm the sample shows the broad band which is due to the stretching vibration of surface hydroxyl groups. This is most likely due to reabsorption from the ambient atmosphere. The peak appeared at 1624.29/cm and 620.75/cm gives the bending mode of OH bonds.

**III. CONCLUSIONS**

In this paper, copper,tin oxide and copper dopedtinoxide nanostructures were synthesized by the sol-gel method. The structural and optical properties using XRD and FTIR were studied. Further, there is a scope of study of other properties like conducting, other morphological properties andthe sensing applications.

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